Avogadro's Law:

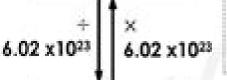
under the same temperature and pressure, some volume of gases will consist of the same number of moles of gases.

Volume ratio = Mole ratio of gases

E.g. if 3 motes of gos A occupies 12 amil, then 3 motes of gos 8 at the same temperature and pressure will also occupy 12 dmil.

number

of particles



% 的复数订字 =

mass of PURE X 100%

 $+M_{\odot}$

 $\times M$

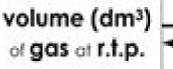
+ M.

% YIELD =

CALCULATED mass × 100%

Limiting reagent: creat when recal most of alternations are given

- Grt. 0.200 ml of H_i was reacted with 8,430 mal of O_i to produce water. Determine the limiting reagent of this reaction.
- Am: 8,200 mol of N. required 8,100 mol of Ov for complete reaction. Nowever 9,400 mol of Ov was supplied. Prenefore: Ox is in excess and his is the limiting reagent.



+ 24 dm³ × 24 dm³

MOLES

×M, mass (g)

of substance

Dilution:

number of moles stays the same, total valume of solution, increases, concentration of solution decreases

conc. (diluted) = conc. (original)
$$\times \frac{\text{wol.(original)}}{\text{xol.(new)}}$$

Sampling:

number of males decreases, valume of solution decreases, concentration remains the same + vol (dm³) of solution

× vol (dm³) of solution

conc. of solution (mol dm⁻³) vol (dm³) of solution

lm³) vol (dm²) tion of solution

conc. of solution

(g dm⁻³)

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Ranjeet Shahi

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