

Part I - Multiple Choice (45 pts, 3 pts each).

1) Which one of the following is not a state function?

- A) enthalpy H
 B) heat q
 C) temperature T
 D) pressure P
 E) internal energy U

Use the table of specific heats to answer questions 2 and 3

2) Which of the following substances would need the largest amount of heat to change the temperature of 1 gram by 10°C?

- A) Aluminum, Al
 B) Copper, Cu
 C) Gold, Au
 D) Lead, Pb
 E) Silver, Ag

Substance	Specific Heat
Aluminum	0.900 J / g °C
Copper	0.386 J / g °C
Gold	0.126 J / g °C
Lead	0.128 J / g °C
Silver	0.233 J / g °C

$$q = cm\Delta T$$

10 g of

3) Which would need the largest amount of heat to change the temperature of 1 mole by 10°C?

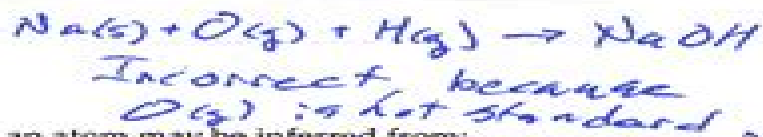
- A) Aluminum, Al
 B) Copper, Cu
 C) Gold, Au
 D) Lead, Pb
 E) Silver, Ag

$$q = C_m \Delta T = M C_m \Delta T$$

10 mole

4) The enthalpy of formation, ΔH_f° , of NaOH (s) is the same as the ΔH_f° for which of these reactions?

- A) $\text{Na (s)} + \text{H}_2\text{O (l)} \rightarrow \text{NaOH (s)} + \frac{1}{2} \text{H}_2 \text{ (g)}$
 B) $\text{Na (s)} + \frac{1}{2} \text{O}_2 \text{ (g)} + \frac{1}{2} \text{H}_2 \text{ (g)} \rightarrow \text{NaOH (s)}$
 C) $\text{Na (s)} + \frac{1}{2} \text{H}_2\text{O}_2 \text{ (l)} \rightarrow \text{NaOH (s)}$
 D) $\text{Na}^+ \text{ (aq)} + \text{OH}^- \text{ (aq)} \rightarrow \text{NaOH (s)}$
 E) $\text{N (g)} + \text{a} + \text{O (g)} + \text{H}^- \text{ (aq)} \rightarrow \text{NaOH (s)}$



5) The existence of discrete (quantized) energy levels in an atom may be inferred from:

- A) diffraction of electrons by crystals
 B) atomic line spectra
 C) visible spectrum
 D) experiments on the photoelectric effect
 E) X-ray diffraction by crystals

6) Arrange the following in order of increasing energy:

Ultra-violet (UV) radiation, visible radiation, gamma-rays, and infra-red (IR) radiation.

IR
 Lowest

Visible

UV
 →

gamma
 Highest

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Rachel S Tattersall



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