



Latitude and Longitude

Longitudes Latitudes Throughout World

Soumitra Kapur



Longitudes Latitudes Throughout World:

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LONGITUDE 155.71°W R.K. Lindsey Jr.,2013 Latitude 20 04 N Longitude 155 71 W is a Hawaiian story told through snapshots and snippets of the lives of fabled and fictionalized family members named Lindsey Beckley and others Through them I recall bits and pieces of Hawai'i's rich and glorious past everyday folks living ordinary lives doing extraordinary things I hope you as the reader will be able to relate to and find the tie that binds us as God's children one planet one people one purpose A wide swath of time is covered in this book of short chapters from the creation of the Hawaiian Islands eighty million years ago to the landing of our people at Wai'ahukini AD 300 to the arrival of James Cook 1778 and the first New England Missionary Company 1820 to the establishment of sugar and pineapple plantations and cattle empires 1800s and to the political trade winds that blew across the archipelago and from chiefdoms to unified kingdom 1819 to provisional government 1893 to republic 1895 to territory 1900 and to statehood 1959 Holly Birch and Mister John are focal characters Latitude 20 04 N Longitude 155 71 W is history written in shorthand in a fun interesting and dynamic way I started this journey with that intent I hope I have achieved my intention You be the judge Always with aloha Bob Lindsey **The World-wide Encyclopedia and Gazetteer** William Harrison De Puy,1899

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web overall the experiment succeeded in showing that temperature and pressure for an ideal gas at constant volume and mass follow the relation of the ideal gas equation differences existed in the experimental graph of temperature versus and pressure and the theoretical curve of temperature versus pressure

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web we then changed the volume between 5ml and 20ml and recorded the pressure at each this part of the lab showed that volume and pressure are inversely proportionate as the pressure would decrease if the volume was increased part b of the experiment looked at the relationship between temperature and pressure

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web procedure 1 the relationship between pressure and volume set the plunger of the syringe to 1 10 the capacity of the syringe the 20 ml mark if using a 20 ml syringe the volume is read from the edge of the bottom rib of the plunger attach the syringe to the connector at the end of sensor

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web he concluded that at a constant pressure the volume of the gas would expand at a particular rate for each degree of temperature rise that being 1 267 per degree in 1874 victor regnault obtained better experimental results showing this number to be 1 273 and concluded that the pressure would approach zero at 1 273 degrees c

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web the volume of 1 00 mol 1 00 mol of any gas at stp standard temperature 273 15 k and pressure 1 atm is measured to be 22 414 l 22 414 l we can substitute 101 325 kpa 101 325 kpa for pressure 22 414 l 22 414 l for volume and 273 15 k 273 15 k for temperature into the ideal gas equation and solve for r r

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web chemists sometimes make comparisons against a standard temperature and pressure stp for reporting properties of gases 273 15 k and 1 atm 101 325 kpa 1 at stp one mole of an ideal gas has a volume of about 22 4 l this is referred to as the standard molar volume figure 9 18

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web figure pageindex 10 since the number of moles in a given volume of gas varies with pressure and temperature changes chemists use standard temperature and pressure 273 15 k and 1 atm or 101 325 kpa to report properties of gases

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web pressure increases even though there were errors in the experiment the goal of the experiment was met the goal of declaring what type of relationship it is was clear using the data pressure kpa temperature c temperature k constant k p t 95 82 kpa 16 0 c 289 k 0 33 97 90 kpa 19 0 c 292 k 0 34 98 97 kpa 99 22 kpa 101 4 kpa 19 1 c

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web figure pageindex 5 the empirically determined relationships among pressure volume temperature and amount of a gas the thermometer and pressure gauge indicate the temperature and the pressure qualitatively the level in the flask indicates the volume and the number of particles in each flask indicates relative amounts

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