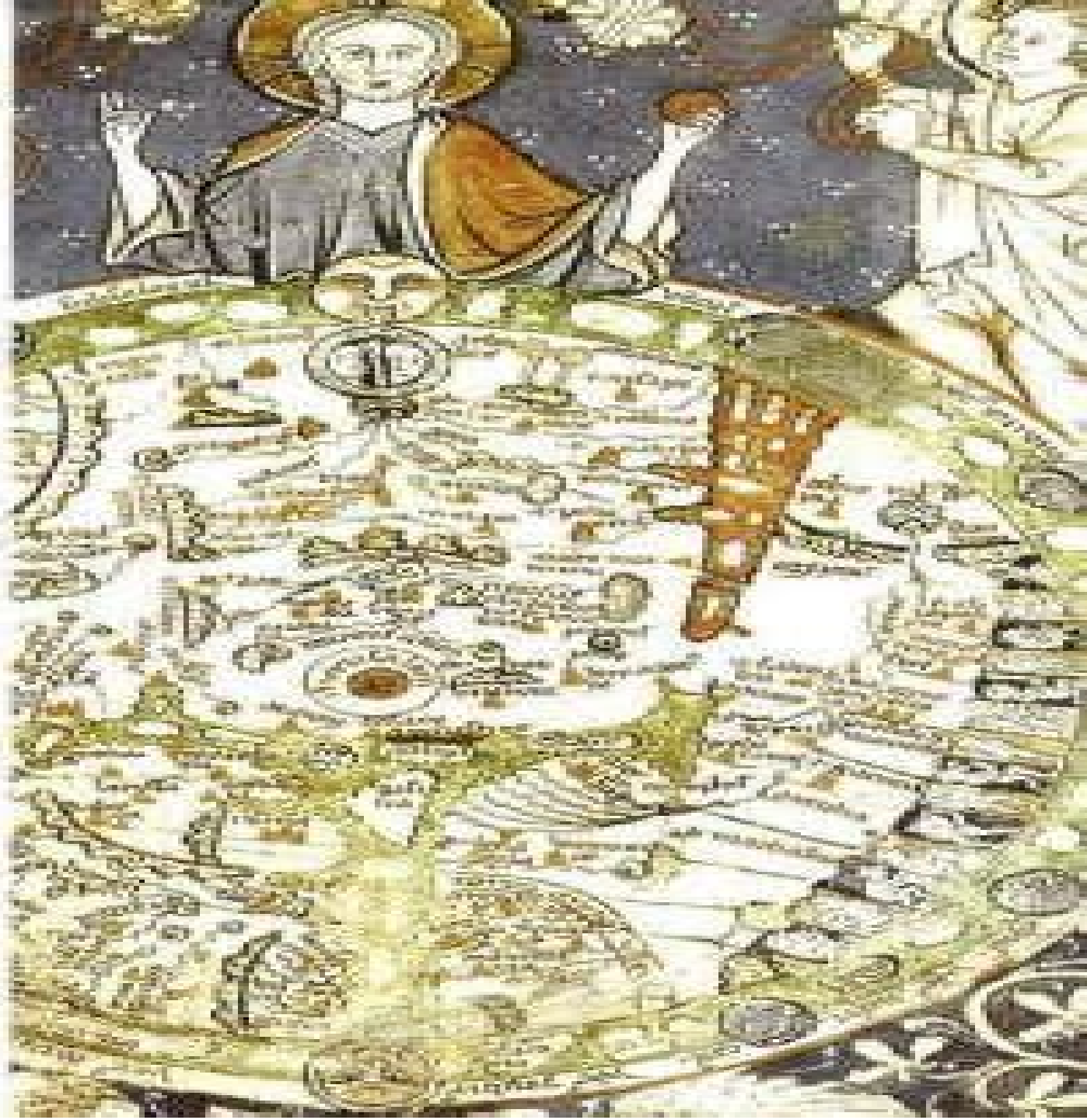


MONSTERS MARVELS AND MIRACLES

EDITED BY LEIF SØNDERGAARD & RASMUS THORNING HANSEN

Imaginary Journeys and Landscapes in the Middle Ages



Monsters Marvels And Miracles

JL Elias



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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3\text{ mole H}_2\text{S} \times \frac{2\text{ mole H}_2\text{O}}{2\text{ mole H}_2\text{S}} \times 18\text{ g H}_2\text{O} = 149.4\text{ g H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1\text{ g H}_2\text{O}}{149.4\text{ g H}_2\text{O}} \times 100 = 91.77\%$ problems 6-8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH_3 can be produced theoretically from the reaction of 5.0 g of N_2 what is the limiting reagent if 8.52 g are actually formed what is the

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al $\times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{3 \text{ mol Al}} \times 101.96 \text{ g/mol} = 63.4 \text{ g}$ c determine the percent yield $\frac{46.0 \text{ g}}{63.4 \text{ g}} \times 100 = 72.6\%$

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