

# Words

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## New Words For You

**Angie Papple Johnston, Rod Powers**



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temperature rose from 23.42 °C to 27.64 °C and the heat capacity of the calorimeter is 4.90 kJ/°C then determine the heat of combustion of sucrose  $C_{12}H_{22}O_{11}$  in kJ per mole of  $C_{12}H_{22}O_{11}$

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web in a calorimeter the white solid  $CaSO_4$  forms and the temperature of the mixture increases to 28.1 °C assuming that the calorimeter absorbs only a negligible quantity of heat and that the specific heat capacity of the solution is 4.18 J/°C·g and that the density of the final solution is 1.0 g/mL calculate the enthalpy change per mole

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web problem 8 5 1 6 8 5 1 6 when 50.0 g of 0.200 M NaCl(aq) at 24.1 °C is added to 100.0 g of 0.100 M AgNO<sub>3</sub>(aq) at 24.1 °C in a calorimeter the temperature increases to 25.2 °C as AgCl(s) forms assuming the specific heat of the solution and products is 4.20 J/g °C calculate the approximate amount of heat in joules produced answer