



The Sable-Hunter.

# Sable Hunter

**Peipei Pang**



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web aim estimation of amino acids by titration methods principle amino acids react with formalin to form methylene amino acids but it was found that the reaction was more complex and the main product is monomethylol and di methylol amino acids amino acids exist in zwitter ionic form and cannot be titrated directly with alkali

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web the sørensen formol titration sft invented by s p l sørensen in 1907 1 is a titration of an amino acid with potassium hydroxide in the presence of formaldehyde 2 it is used in the determination of protein content in samples 3 formol titration equation for amino acids in general

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web the addition of formaldehyde to amino acids and the resulting increase in their acidity towards phenol phthalein or thymol phthalein constitutes the basis of the well known sørensen formol method 2 for their estimation by titration with soda accordingly some knowledge of the acid base

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web amino acids formol titration an alternative simpler procedure for improving the inflexion in the neutralization of an amino acid is to add formaldehyde to the solution although this does not affect the acid titration curve the one for

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web the principle of this method is that formaldehyde make blocking of amino group of amino acid so amino acid become

more acidic monomethylol and dimethylol derivative of amino acid and titrated

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web presence of a large excess formaldehyde amino acid can be titrated to a phenolphthalein end point when amino acid reacts with formaldehyde  $\text{NH}_2$  amino group in the amino acid will react with formaldehyde to form methylol and dimethylol derivatives which in

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web the amino acid solutions were at approximately pH 6 prior to the addition of formaldehyde the end point of the titration was at about pH 9 no correction was made for the volume of standard base required to bring the aqueous formaldehyde solution to the end point pH of the amino acid titration and the concentration of the formaldehyde

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web objectives after completing this section you should be able to draw the predominant form of a given amino acid in a solution of known ph given the isoelectric point of the amino acid describe briefly how a mixture of amino acids may

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