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web the rate determining slower step gives a rate law showing second order dependence on the  $[A]^2$  concentration and the sum of the two elementary equations gives the overall net reaction in general when the rate determining slower step is the first step in the reaction mechanism the rate law for the overall reaction is the same as the rate

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by its rate law for example the rate of the gas phase decomposition of dinitrogen pentoxide  $2\text{N}_2\text{O}_5 \rightarrow 2\text{N}_2 + 5\text{O}_2$  has been found to be directly proportional to the concentration of  $\text{N}_2\text{O}_5$  text rate  $k$  n

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web the rate determining step limits the overall rate and therefore determines the rate law for the overall reaction

introduction multistep reactions one of the most important applications of kinetics is to the study of reaction mechanisms or the sequences of steps by which chemical reactions occur

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web the initial rate law is then measured for each of the reactions consider the reaction between nitrogen monoxide gas and hydrogen gas to form nitrogen gas and water vapor the following data were collected for this reaction at 1280 text o text c

see table below notice that the starting concentrations of  $\text{ce no}$  and  $\text{ce h}$

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web determination of a rate law and activation energy in this experiment glass beakers pipettes and a hotplate were utilized to set up the procedures a solution of deionized water a buffer potassium iodide  $\text{ki}$  sodium thiosulfate  $\text{na}_2\text{s}_2\text{o}_3$  starch and hydrogen peroxide  $\text{h}_2\text{o}_2$  were used in order to calculate the rate of reaction

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