

Part I - Multiple Choice (45 pts, 3 pts each).

1) Which one of the following is not a state function?

- A) enthalpy H
 B) heat q
 C) temperature T
 D) pressure P
 E) internal energy U

Use the table of specific heats to answer questions 2 and 3

2) Which of the following substances would need the largest amount of heat to change the temperature of 1 gram by 10°C ?

- A) Aluminum, Al
 B) Copper, Cu
 C) Gold, Au
 D) Lead, Pb
 E) Silver, Ag

| Substance | Specific Heat |
|-----------|-----------------------------------|
| Aluminum | $0.900 \text{ J/g}^\circ\text{C}$ |
| Copper | $0.386 \text{ J/g}^\circ\text{C}$ |
| Gold | $0.126 \text{ J/g}^\circ\text{C}$ |
| Lead | $0.128 \text{ J/g}^\circ\text{C}$ |
| Silver | $0.233 \text{ J/g}^\circ\text{C}$ |

$$q = cm\Delta T$$

\downarrow
10 g
 \downarrow
10 $^\circ\text{C}$

3) Which would need the largest amount of heat to change the temperature of 1 mole by 10°C ?

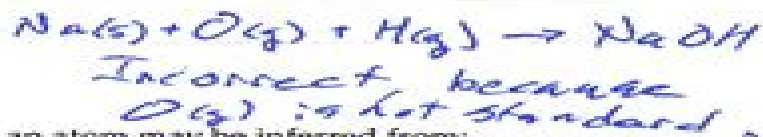
- A) Aluminum, Al
 B) Copper, Cu
 C) Gold, Au
 D) Lead, Pb
 E) Silver, Ag

$$q = C_m \Delta T = M C_m \Delta T$$

\downarrow
10 mole $^\circ\text{C}$

4) The enthalpy of formation, ΔH_f° , of NaOH (s) is the same as the $\Delta H^\circ_{\text{rxn}}$ for which of these reactions?

- A) $\text{Na (s)} + \text{H}_2\text{O (l)} \rightarrow \text{NaOH (s)} + \frac{1}{2} \text{H}_2 \text{ (g)}$
 B) $\text{Na (s)} + \frac{1}{2} \text{O}_2 \text{ (g)} + \frac{1}{2} \text{H}_2 \text{ (g)} \rightarrow \text{NaOH (s)}$
 C) $\text{Na (s)} + \frac{1}{2} \text{H}_2\text{O}_2 \text{ (l)} \rightarrow \text{NaOH (s)}$
 D) $\text{Na}^+ \text{ (aq)} + \text{OH}^- \text{ (aq)} \rightarrow \text{NaOH (s)}$
 E) $\text{N (g)} + \text{a} + \text{O (g)} + \text{H}^- \text{ (aq)} \rightarrow \text{NaOH (s)}$



5) The existence of discrete (quantized) energy levels in an atom may be inferred from:

- A) diffraction of electrons by crystals
 B) atomic line spectra
 C) visible spectrum
 D) experiments on the photoelectric effect
 E) X-ray diffraction by crystals

6) Arrange the following in order of increasing energy:

Ultra-violet (UV) radiation, visible radiation, gamma-rays, and infra-red (IR) radiation.

IR visible UV gamma
 Lowest Highest

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**Robert Czajkowski, Krishna Mohan
Poluri**



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Engineering Synthetic Metabolons: From Metabolic Modelling to Rational Design of Biosynthetic Devices Lars M. Voll,Zoran Nikoloski,2016-07-19 The discipline of Synthetic Biology has recently emerged at the interface of biology and engineering The definition of Synthetic Biology has been dynamic over time ever since which exemplifies that the field is rapidly moving and comprises a broad range of research areas In the frame of this Research Topic we focus on Synthetic Biology approaches that aim at rearranging biological parts entities in order to generate novel biochemical functions with inherent metabolic activity This Research Topic encompasses Pathway Engineering in living systems as well as the in vitro

assembly of biomolecules into nano and microscale bioreactors Both the engineering of metabolic pathways in vivo as well as the conceptualization of bioreactors in vitro require rational design of assembled synthetic pathways and depend on careful selection of individual biological functions and their optimization Mathematical modelling has proven to be a powerful tool in predicting metabolic flux in living and artificial systems although modelling approaches have to cope with a limitation in experimentally verified reliable input variables This Research Topic puts special emphasis on the vital role of modelling approaches for Synthetic Biology i e the predictive power of mathematical simulations for i the manipulation of existing pathways and ii the establishment of novel pathways in vivo as well as iii the translation of model predictions into the design of synthetic assemblies

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