



The New Official LSAT
TriplePrep
Volume 2™

PrepTests 155–157

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RD Boyd



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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} \rightarrow \text{Au}(\text{CN})_2^- + 2\text{e}^-$ $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ make electrons equal

convert to basic solution 4cn 2au 2au cn 2 2e multiplied by a factor of 2 2e 2h₂o o₂ h₂o₂ 2oh 3 the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s ni a 2 a q 2 e now that the oxidation half reaction is balanced it tells us that two electrons are produced for

every atom of nickel oxidized but where do those electrons go

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6 + 2\text{HNO}_2 + 2\text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} \rightarrow 2\text{HNO}_2 + 2\text{C}_6\text{H}_8\text{O}_6 + 2\text{H}_2\text{O}$ choose 1 answer c $\text{C}_6\text{H}_8\text{O}_6$ and HNO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and H_2O b HNO_2 and $\text{C}_6\text{H}_8\text{O}_6$ h and H_2O c h and