

Date Performed: 10-17-05

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Acid-Base Titrations

Objective:

- ❖ To standardize a sodium hydroxide solution by its reaction with a hydrochloric acid solution of known concentration.
- ❖ To determine the molar mass of an unknown, monoprotic acid.
- ❖ To determine the percentage of acetylsalicylic acid in an aspirin tablet

Experimental Data

In Experiment A)

Molarity of HCl 0.200M

NaOH used in first titration 20.78mL

NaOH used in second titration 20.80mL

In Experiment B)

NaOH used in first titration 20.48 mL

NaOH used in second titration 20.40 mL

In Experiment C)

NaOH used to titrate aspirin 7.15 mL

Sample Calculations

Part A

$$\begin{aligned}\text{Mol NaOH} &= (M_{\text{HCl}})(V_{\text{HCl}}) = \text{Mol HCl} \\ &= (0.200\text{M})(0.02500\text{L}) \\ &= 0.00500 \text{ mol}_{\text{HCl}} = \text{mol}_{\text{NaOH}}\end{aligned}$$

Molarity (M) = Moles of solute / Volume of solution in liters

$$\begin{aligned}(M)_{\text{NaOH}} &= \text{Moles}_{\text{NaOH}} / V_{\text{NaOH}} \\ &= 0.500\text{mol}_{\text{NaOH}} / 0.02080\text{L} \\ &= 0.240\text{M}_{\text{NaOH}}\end{aligned}$$

Part B

$$\begin{aligned}(M_{\text{NaOH}})(V_{\text{NaOH}}) &= \text{Moles}_{\text{unknown}} \\ (0.240\text{M})(0.02048\text{L}) &= 0.00492\text{mol} \\ \text{Molar Mass} &= \text{Mass of acid} / n_{\text{A}}, \text{Moles of acid} \\ \text{Molar Mass}_{\text{unknown}} &= 1.00\text{g} / 0.00492\text{mol} \\ &= 203\text{g/mol}\end{aligned}$$

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