

WS grade 12 (17.3)

- 1- 1M of nitrogen gas and 1M of hydrogen gas placed in a closed container of volume 2L. Calculate the concentrations of all substances at equilibrium knowing that the concentration of hydrogen gas was found 0.2M at equilibrium, then deduce the value of equilibrium constant.

$$[N_2]_0 = 1M$$

$$[H_2]_0 = 1M$$

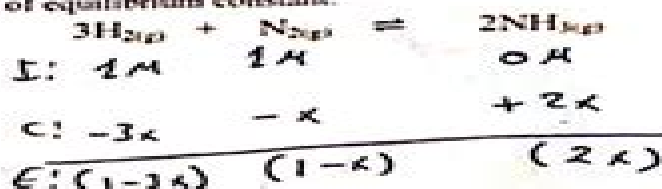
$$V = 2L$$

$$[H_2]_{eq} = 0.2M$$

$$[N_2]_{eq} = ?$$

$$[NH_3]_{eq} = ?$$

$$K_{eq} = ?$$



$$[H_2]_{eq} = 0.2 = 1 - 3x \Rightarrow x = 0.26$$

$$[N_2]_{eq} = 1 - x = 1 - 0.26 = \boxed{0.74M}$$

$$[NH_3]_{eq} = 2x = 2(0.26) = \boxed{0.52M}$$

$$K_{eq} = \frac{[NH_3]^2}{[H_2]^3[N_2]} = \frac{(0.52)^2}{(0.2)^3(0.74)} = \boxed{45.67}$$

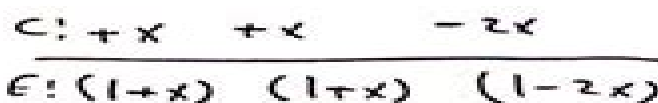
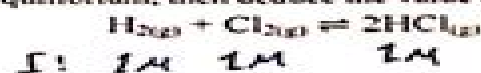
- 2- 1M of HCl, Cl₂, and H₂ placed in a closed container. Calculate the concentrations of all substances at equilibrium knowing that the concentration of hydrogen gas was found 1.3M at equilibrium, then deduce the value of equilibrium constant.

$$[HCl]_0 = 1M$$

$$[Cl_2]_0 = 1M$$

$$[H_2]_0 = 1M$$

$$[H_2]_{eq} = 1.3M$$



$$[H_2]_{eq} = 1.3 = 1 + x \Rightarrow x = 0.3$$

$$[Cl_2]_{eq} = 1 + x = \boxed{1.3M}$$

$$[HCl]_{eq} = 1 - 2(0.3) = \boxed{0.4M}$$

$$K_{eq} = \frac{[HCl]^2}{[H_2][Cl_2]} = \frac{(0.4)^2}{(1.3)(1.3)} = \boxed{0.094}$$

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1. Understanding the eBook Selected Exercises In Chemistry
 - The Rise of Digital Reading Selected Exercises In Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Selected Exercises In Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Selected Exercises In Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Selected Exercises In Chemistry
 - Personalized Recommendations
 - Selected Exercises In Chemistry User Reviews and Ratings
 - Selected Exercises In Chemistry and Bestseller Lists
5. Accessing Selected Exercises In Chemistry Free and Paid eBooks
 - Selected Exercises In Chemistry Public Domain eBooks
 - Selected Exercises In Chemistry eBook Subscription Services

- Selected Exercises In Chemistry Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Selected Exercises In Chemistry Compatibility with Devices
 - Selected Exercises In Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Selected Exercises In Chemistry
 - Highlighting and Note-Taking Selected Exercises In Chemistry
 - Interactive Elements Selected Exercises In Chemistry
- 8. Staying Engaged with Selected Exercises In Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Selected Exercises In Chemistry
- 9. Balancing eBooks and Physical Books Selected Exercises In Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Selected Exercises In Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Selected Exercises In Chemistry
 - Setting Reading Goals Selected Exercises In Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Selected Exercises In Chemistry
 - Fact-Checking eBook Content of Selected Exercises In Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

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- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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