Chemistry CT-1 (Atomic Structure, Moles, Chemical Bonding and Acid-Base-Salt) Q1.

Solid L is a mixture of two compounds. The two compounds contain the same cation but different anions.

- (a) An excess of dilute hydrochloric acid is added to L. Bubbles of carbon dioxide gas are given off and the solid dissolves completely, forming a colorless solution.
 - (i) What conclusion can be made from the fact that the solution is colorless?

 [1]
 - (ii) Give a test and observation to identify carbon dioxide gas.

observation [2]

(iii) Identify the anion which reacts with hydrochloric acid to produce carbon dioxide gas.

[1]

The colorless solution formed in (a) is divided into two parts for tests (b) and (c). Complete the table.

test		observation	conclusion	
(b)(i)	To the first part, in a test-tube, aqueous sodium hydroxide is added until a change is seen.	white precipitate		
(H)	An excess of aqueous sodium hydroxide is added to the mixture from (i).	insoluble in excess		
(c)			L contains NO ₃ ions.	[1
				[4

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