

MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

- 1) (3pts) What is the conjugate acid of NH_3 ?
 A) NH_3 B) NH_4OH C) NH_4^+ D) NH_2^+ E) NH_3^+
- 2) (3pts) In basic solution, _____.
 A) $[\text{OH}^-] > 7.00$
 B) $[\text{H}_3\text{O}^+] > [\text{OH}^-]$
 C) $[\text{H}_3\text{O}^+] = 0 \text{ M}$
 D) $[\text{H}_3\text{O}^+] = [\text{OH}^-]$
 E) $[\text{H}_3\text{O}^+] < [\text{OH}^-]$
- 3) (3pts) What is the pH of an aqueous solution at 25° that contains $3.98 \times 10^{-9} \text{ M}$ hydronium ion?
 A) 3.980 B) 8.400 C) 7.000 D) 5.600 E) 9.000
- 4) (3pts) Which solution below has the highest concentration of hydroxide ions?
 A) $\text{pH} = 7.93$ B) $\text{pH} = 3.21$ C) $\text{pH} = 7.00$ D) $\text{pH} = 12.59$ E) $\text{pH} = 9.82$
- 5) (3pts) The hydride ion, H^- , is a stronger base than the hydroxide ion, OH^- . The product(s) of the reaction of hydride ion with water is/are _____.
 A) no reaction occurs
 B) $\text{H}_3\text{O}^+ (\text{aq})$
 C) $\text{H}_2\text{O}_2 (\text{aq})$
 D) $\text{OH}^- (\text{aq}) + \text{H}_2 (\text{g})$
 E) $\text{OH}^- (\text{aq}) + 2\text{H}^+ (\text{aq})$
- 6) (3pts) Of the acids in the table below, _____ is the strongest acid.

Acid	K_a
HOAc	1.8×10^{-5}
HCHO_2	1.8×10^{-4}
HClO	3.0×10^{-8}
HF	6.8×10^{-4}

- A) HF
- B) HCHO_2
- C) HOAc and HCHO_2
- D) HOAc
- E) HClO

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Markow



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Table of Contents Pkg Acp Chem 122

1. Understanding the eBook Pkg Acp Chem 122
 - The Rise of Digital Reading Pkg Acp Chem 122
 - Advantages of eBooks Over Traditional Books
2. Identifying Pkg Acp Chem 122
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pkg Acp Chem 122
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pkg Acp Chem 122
 - Personalized Recommendations
 - Pkg Acp Chem 122 User Reviews and Ratings
 - Pkg Acp Chem 122 and Bestseller Lists
5. Accessing Pkg Acp Chem 122 Free and Paid eBooks
 - Pkg Acp Chem 122 Public Domain eBooks
 - Pkg Acp Chem 122 eBook Subscription Services

- Pkg Acp Chem 122 Budget-Friendly Options
- 6. Navigating Pkg Acp Chem 122 eBook Formats
 - ePub, PDF, MOBI, and More
 - Pkg Acp Chem 122 Compatibility with Devices
 - Pkg Acp Chem 122 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pkg Acp Chem 122
 - Highlighting and Note-Taking Pkg Acp Chem 122
 - Interactive Elements Pkg Acp Chem 122
- 8. Staying Engaged with Pkg Acp Chem 122
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pkg Acp Chem 122
- 9. Balancing eBooks and Physical Books Pkg Acp Chem 122
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pkg Acp Chem 122
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pkg Acp Chem 122
 - Setting Reading Goals Pkg Acp Chem 122
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pkg Acp Chem 122
 - Fact-Checking eBook Content of Pkg Acp Chem 122
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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