

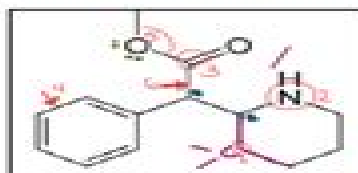
Chem 104 Kelly's CLC Exam 1 Prep & Review

Module 1 Organic Chem Learning Outcomes and Related Student Skills

1.1 Draw and evaluate representations of organic molecules

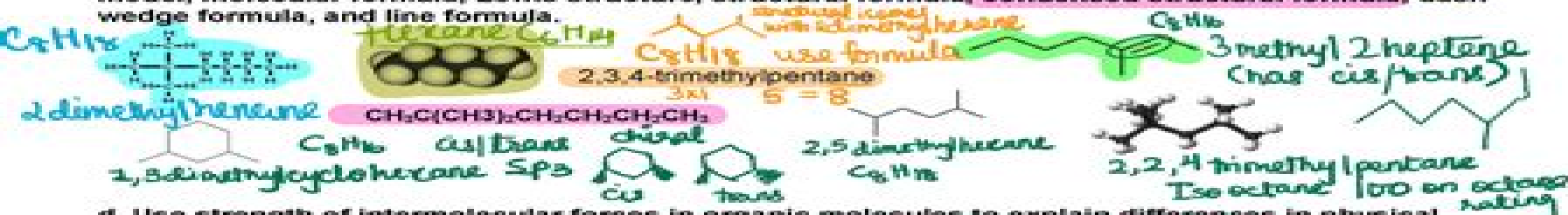
- a. Identify the molecular geometry of a given atom and the bond angles in organic compounds.
b. Identify hybrid orbitals used in bonding and use hybridization of carbon atoms (sp , sp^2 , and sp^3) to rationalize molecular structure.

Total Regions	Bonding Regions	Lone Pair Regions	Electron Pair Geometry	Molecular Geometry	Hybridization of central atom	Bond Angle
2	2	0	Linear	Linear	sp	180°
3	3	0	Trigonal Planar	Trigonal Planar	sp ²	120°
3	2	1	Trigonal Planar	Bent	sp ²	<120°
4	4	0	Tetrahedral	Tetrahedral	sp ³	109.5°
4	3	1	Tetrahedral	Trigonal Pyramidal	sp ³	<109.5°
④	③	①	Tetrahedral	Bent	sp ³	<109.5°



1. What's the bond angle here? Sp^3 , tetrahedral 109.5°
 2. What's the molecular geometry here? Sp^3 , trigonal pyramidal ($< 109.5^\circ$)
 3. What's the hybridization of this carbon? Sp^2 trigonal planar 120°
 4. Describe the bond in terms of orbital overlap. Also tell the bond order $\sigma C_1Sp^2 + C_2Sp^2$
 5. What's the bond angle here? $< 109.5^\circ$ Bent Sp^3 $\pi C_12p + C_22p$
 6. What orbitals overlap to make this bond? Sigma C_1sp^3
- they need to be Sp^3 hybridized to be linear
- Methylidenecyclopropane

c. Use and understand different representations of molecules: ball and stick model, space-filling model, molecular formula, Lewis structure, structural formula, condensed structural formula, dash-wedge formula, and line formula.



d. Use strength of intermolecular forces in organic molecules to explain differences in physical properties such as boiling point and solubility. *increase in decreasing BP and solubility*

1. Compare the boiling points of hexane, butanoic acid and methylpropanoate. Draw, explain.



e. Given the molecular formula of an organic compound, draw all of its structural isomers.

Draw and name all isomers of $C_3H_5O_2$ that don't contain a ring. *Draw four more*



f. Determine whether a molecule has optical isomers (enantiomers) and identify chiral carbon atoms.

1. What is a chiral carbon? Sp³ hybridised 4 distinct different attached groups (Y, not V, X)
2. Which of the molecules in 1.1c (above) have any chiral carbons?
3. What is an enantiomer?
(non-superimposable) mirror image contain chiral carbon
optical isomers - rotate polarised light

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K Morrison



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