

Fig. 2.1 shows the general configuration of a thermodynamic system at equilibrium at temperature T . The internal energy of the gas is the important macroscopic (state) variable.

Assume that the walls of the cylinder are rigid and adiabatic, so that

$$\frac{dQ}{dt} = \frac{dU}{dt} = \frac{dW}{dt}$$

where U is the internal energy, Q is the heat, W is the work done, and T is the temperature. The walls are rigid and adiabatic, so that

where

$$P_{ext} = \frac{F}{A} = \frac{\text{force exerted by the gas}}{\text{area of the piston}} = P_{int} \text{ is usually}$$

Fig. 2.2 shows the general configuration of a thermodynamic system at equilibrium at temperature T . The internal energy of the gas is the important macroscopic (state) variable. The walls are rigid and adiabatic, so that

$$\frac{dQ}{dt} = \frac{dU}{dt} = \frac{dW}{dt}$$

$$\text{where } P_{ext} = \frac{\text{force exerted by the gas}}{\text{area of the piston}} = P_{int} \text{ is usually}$$

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