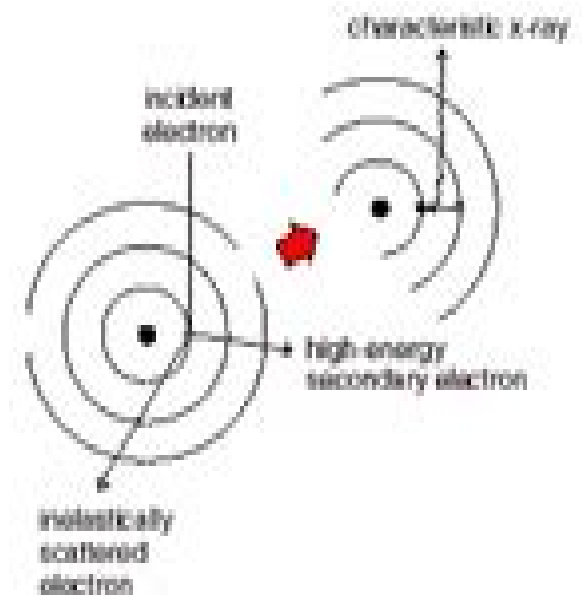


K-shell Emission Radiation

- Remember that atoms have their electrons arranged in closed “shells” of different energies. The K-shell is the lowest energy state of an atom.
- An incoming electron can give a K-shell electron enough energy to knock it out of its energy state.
- About 0.1% of the electrons produce K-shell vacancies; most produce heat. Then, a tungsten electron of higher energy (from an outer shell) can fall into the K-shell.
- The energy lost by the falling electron shows up as an emitted X-ray photon. Meanwhile, higher energy electrons fall into the vacated energy state in the outer shell, and so on.
- After losing an electron, an atom remains ionized for a very short time (about 10^{-14} second) and thus an atom can be repeatedly ionized by the incident electrons which arrive about every 10^{-12} second.
- Generally, K-shell emission produces higher-intensity X-rays than Bremsstrahlung, and the X-ray photon comes out at a single wavelength.



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represented as graphs comprised of nodes neuronal elements and edges their connections and advanced signal processing techniques **Subject Guide to Books in Print** ,1993 **The Soviet Journal of Nondestructive Testing** ,1967

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