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# Long Road Home One Step At A Time

**R Pring**



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web the molar volume will be the volume occupied by one molar mass 32 g of the liquid expressing the density in liters instead of ml we have  $v = \frac{m}{\rho}$  32 g mol / 1.790 g l = 17.87 l mol

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web the molar volume has the SI unit of cubic metres per mole  $\text{m}^3 \text{mol}^{-1}$  although it is more typical to use the units cubic decimetres per mole  $\text{dm}^3 \text{mol}^{-1}$  for gases and cubic centimetres per mole  $\text{cm}^3 \text{mol}^{-1}$  for liquids and solids

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web zero degrees celsius or 273 K is defined as standard temperature and one atmosphere or 760 mm Hg is defined as standard pressure the volume occupied by one mole of a gas is called molar volume therefore the molar volume of a gas at standard temperature and pressure STP is  $22.4 \text{ L}$

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relationships between these properties are explained at a molecular level and simple models are introduced that allow the various gas laws to be derived from first principles finally the collisional behavior of gases is used to

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web jun 8 2023  $V = 1.00 \text{ mol} \times 0.08206 \text{ L atm mol}^{-1} \text{ K}^{-1} \times 273 \text{ K} = 22.4 \text{ L}$  in other words 1 mole of a gas will occupy 22.4 l at stp assuming ideal gas behavior at stp the volume of a gas is only dependent on number of moles of

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