



# S Dumoi O Narodnom Prosveshchenii Iz Perepiski Chast 2

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web log  $[\text{H}^+]$   $10^{-4.05}$   $8.91 \times 10^{-5}$  mol dm<sup>-3</sup> log  $\text{pK}_a$   $\text{pH}$  3.75 4.05 0.30 10.0 3.0 0.50 and  $\text{HCOOH}$  0.1 so  $\text{HCOO}^-$  0.200 mol dm<sup>-3</sup> amount  $\text{H}^+$   $\text{OH}^-$

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web only the better candidates could write successive ionisation equations in part c while part d about the  $\text{pH}$  of equimolar strong and weak acid solutions was poorly answered

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web section 2  $\text{pH}$  of strong acids number of protons released monoprotic acid acid that releases one  $\text{H}^+$  ion per molecule e.g.  $\text{HCl}$  hydrochloric acid  $\text{HNO}_3$  nitric acid  $\text{CH}_3\text{COOH}$  acetic acid

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