

Bond-order and Magnetic Behavior of Diatomic Species without Molecular Orbital Theory

Arijit Das*

Department of Chemistry, Ramthakur College, Agartala -799003, Tripura West, Tripura, India

*Corresponding author: arijidas78chem@gmail.com

Abstract In this chapter text based learning approaches has been highlighted by innovative and time economic way to enhance interest of students' who belong to paranoia zone in Electronic Structure of Atoms and Molecules beneath Inorganic Chemistry of chemical science. In this pedagogical survey, I have tried to hub two (02) time economic pedagogies by including seven (07) new formulae in the field of chemical education. This chapter explores the results and gives implications for context based teaching, learning and assessment.

Keywords: general public, high school, graduate student, chemical bonding, bond order, paramagnetic, diamagnetic, molecular orbital theory

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1. Introduction

The conventional methods [1-5] for prediction of bond order and magnetic properties of diatomic species having (1-20)e's using molecular orbital theory is time consuming. Keeping this in mind, I have introduced some innovative methods [6-12] in this book chapter to make chemistry metabolic, time economic and interesting.

2. Bond-Order of Diatomic Species without MOT

Bond-order usually predicted from the Molecular Orbital Theory. Molecular Orbital Theory (M.O.T.) was first proposed by Friedrich Hund and Robert Mulliken in

1933. They developed an approach to covalent bond formation which is based upon the effects of the various electron fields upon each other and which employs molecular orbital rather than atomic orbital. Each such orbital characterizing the molecule as a whole is described by a definite combination of quantum numbers and possesses relative energy value.

For homo and hetero nuclear diatomic molecules or ions having (1-20)e's.

The graphical representation presented in Figure 1 shows that bond-order gradually increases to 1 in the range (0-2) electrons then falls to zero in the range (2-4) electrons then it further rises to 1 for (4-6) electrons and once again falls to zero for (6-8) electrons then again rises to 3 in the range (8-14) electrons and then finally falls to zero for (14-20) electrons. For total no of electrons 2, 6 and 14, we may use multiple formulae, because they fall in the overlapping region in which they intersect with each other.

Graphical Representation of B.O. with number of electrons:

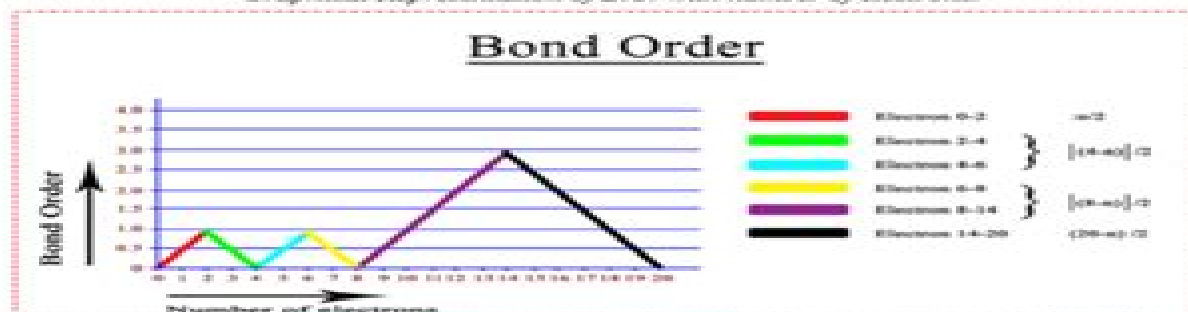


Figure 1. (B.O. vs. number of electrons)

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Evgeny Y. Tsymbal, Igor Žutić



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